**QUIZ Review**

**For #1-4:** Identify the TYPE of reaction and BALANCE the equation. Indicate the state of matter for the products of each reaction.

1. Na2SO4 (aq) + CaCl2 (aq) 🡪 NaCl ( ) + CaSO4 ( )
2. aluminum (s) + copper(II) chloride (aq) 🡪 copper ( ) + aluminum chloride ( )
3. (NH4)2S (aq) + Fe(NO3)3 (aq) 🡪 NH4NO3 ( ) + Fe2S3 ( )
4. C2H6 burns

**For #5-10:** Identify if each of the following WILL or WILL NOT result in a chemical change. Explain **WHY** you know if there is a chemical change or not.

1. zinc + lithium chloride 🡪
2. Ca(OH)2 + HCl 🡪
3. C3H6O + O2 🡪
4. CaI2 + Ni(NO3)2 🡪
5. hydrobromic acid + lead(II) nitrate 🡪
6. calcium + tin(IV) chloride 🡪

**For #11-20:**

 1) *Identify the TYPE of reaction*

*2) Determine IF a reaction will happen or if there is “NR”*

*3) If there is a reaction, predict the products and write all chemical symbols correctly*

*4) Balance the equation*

1. Na + F2 🡪
2. sodium acetate solution is mixed with ammonium phosphate solution 🡪
3. aluminum carbonate decomposes 🡪
4. Br2 + AlCl3 🡪
5. magnesium chlorate decomposes 🡪
6. propane (C3H8) burns
7. sodium is mixed with chromium(III) sulfate solution 🡪
8. Ag2O 🡪
9. Na2CO3 + Al2(SO4)3 🡪
10. silver nitrate solution is mixed with calcium chloride solution 🡪